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Chapter 9 Chemical Names and Formulas SECTION 9.3 NAMING AND WRITING FORMULAS FOR MOLECULAR COMPOUNDS (pages 268–270) This section explains the rules for naming and writing formulas for binary molecular compounds. Naming Binary Molecular Compounds (pages 268–269) 1. Circle the letter of the type(s) of elements that form binary molecular compounds.

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When the anion ends in -ite, the acid name is the stem of the anion with the suffix -ous. Example  $\text{H}_2\text{SO}_3$  (X=Sulfite), Sulfurous acid When the anion name ends in -ate, the acid name is the stem of the anion with the suffix -ic.

### **Chapter 9: Chemical Names and Formulas**

OBJ: 9.1.3 Identify the two common endings for the names of most polyatomic ions. | 9.5.3 Apply the rules for writing chemical formulas by using a flowchart. STA: Ch.2 33.ANS: OBJ: 9.5.3 Apply the rules for writing chemical formulas by using a flowchart.

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Names and Formulas for Bases A base is an ionic compound that produces hydroxide ions (OH<sup>-</sup>) when dissolved in water (the Arrhenius definition) Bases are named the same way as other ionic compounds:  
â€œThe name of the cation (which is a metal) is followed by the name of the anion (which will be hydroxide).

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Example 9.2What is the molecular mass of the compound which has the formula Ca(NO<sub>3</sub>)<sub>2</sub>? It is necessary to understand that the parentheses in chemical formulas group together collections of atoms of constant composition; submolecules if you will. The NO<sub>3</sub> group occurs frequently enough in different chemical compounds to have its own name, nitrate.

### **Chapter 9. Chemical Formulas - University of New Mexico**

Chapter 6 CHEMICAL NAMES AND FORMULAS Section 6.1 INTRODUCTION TO CHEMICAL BONDING Section 6.2 REPRESENTING CHEMICAL COMPOUNDS Section 6.3 IONIC CHARGES Section 6.4 IONIC COMPOUNDS Section 6.5 MOLECULAR COMPOUNDS AND ACIDS Section 6.6 SUMMARY OF NAMING AND FORMULA WRITING FEATURES DISCOVER IT! Element Name Search SMALL-SCALE LAB

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The name of a base is the name of the cation followed by the name of the anion (hydroxide). The formula of a base is written by showing the number of hydroxide ions needed to balance the positive charge on the cation. After reading Lesson 9.4, answer the following questions. Names and Formulas of Acids hydrogen1.

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282 Chapter 9 â€œ Chemical Reactions Section 9.1.1 Figure 9.1 When adipoyl chloride in dichloromethane reacts with hexanediamine, nylon is formed. Nylon is used in many products, including carpeting, clothing, sports equipment, and tires. Objectives Recognize evidence of chemical change. Represent chemical reactions with equations.

## Chapter 9: Chemical Reactions

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CHAPTER 9 Chemical Names And Formulas Test Review. STUDY. PLAY. Monoatomic Ions. Consist of a single atom with a positive or negative charge resulting from the loss of gain of one or more valence electrons, respectively. Cations. Metallic elements that tend to lose valence electrons.

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Chapter 9 Chemical Names and Formulas 221 . 4. Write formulas for the ionic compounds. e. onium chloride f. potassium dichromate a. sodium phosphate b. magnesium sulfate c. sodium hydroxide d. potassium cyanide 5. Write formulas for compounds formed from these pairs of ions. a.  $\text{NH}_4^+$ .  $\text{SO}_4^{2-}$

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Convert between the names of compounds and their chemical formulas. (Section 5.3) Report the answers to calculations to the correct number of significant figures. (Section 8.2) Use percentages as conversion factors. ... 332 Chapter 9 Chemical Calculations and Chemical Formulas.

### Chapter 9 Chemical Formulas and Chemical Compounds - Mark Bishop

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\_\_\_\_ 18. What is the name of H<sub>2</sub>SO<sub>3</sub>? a. hyposulfuric acid c. sulfuric acid b. hydrosulfuric acid d. sulfurous acid \_\_\_\_ 19. When the name of an anion that is part of an acid ends in -ite, the acid name includes the suffix \_\_\_\_\_. a. -ous c. -ate b. -ic d. -ite \_\_\_\_ 20. What is the formula for sulfurous acid? a. H<sub>2</sub>SO<sub>4</sub> c. H<sub>2</sub>SO<sub>2</sub> b. H<sub>2</sub>SO<sub>3</sub> d. H<sub>2</sub>S \_\_\_\_ 21.

### Chapter 9 Practice Test

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### Chapter-09---Chemical-Names-and-Formulas.pdf - CHAPTER

name date class chemical names and formulas 9 - chapter 9 chemical names and formulas85 names and formulas for bases (page 273) 6. a base is a compound that produces \_\_\_\_\_ when dissolved in water.

### Chapter 9 Chemical Names And Formulas PDF

Ch. 5 Notes: TYPES OF COMPOUNDS I. Compounds in General A. compound (â€•cmpd.â€•)â€•a substance formed from more than one element 1) has more than one different capital letterâ€• Al 2 S 3, NaBr, Ca(OH) 2 2) can be separated by chemical means B. moleculeâ€•a group of atoms with no net charge C. two general types of compounds

### Ch. 5 Notes: TYPES OF COMPOUNDS - Kwanga.net

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Chapter 9 "Chemical Names and Formulas" • Click to add text H2O 2 Section 9.1 Naming Ions OBJECTIVES: "Identify the charges on monatomic ions by using the periodic table, and name the ions. 3 Section 9.1 Naming Ions OBJECTIVES: "Define a polyatomic ion and write the names and formulas of the most common polyatomic ions. 4 Section 9.1 ...

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Chapter 9 Chemical Names and Formulas. 233 \_\_\_\_\_ 12. Nitrogen reacts with oxygen to form two compounds. Compound A contains 2.8 g of nitrogen for each 1.6 g of oxygen.

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