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Strong acids and bases ionize to a large extent to form many hydrogen or hydroxide ions and have a large acid or base dissociation constant (K_a or K_b). Weak acids or bases ionize to a small extent to form few hydrogen or hydroxide ions and have a small K_a or K_b . An example of a weak acid is acetic acid.

CHEMISTRY NOTES “ CHAPTERS 20 AND 21 Acids and Bases

Note Taking Acid Bases And Salts Answer Pdf chapter2 acids, bases and salts - acids, bases and salts chapter2 you have learnt in your previous classes that the sour and bitter tastes of food are due to acids and bases, respectively, present in ...

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Note pdf no. 2 What it covers? Bronsted Lowry Concept of Acid and base, conjugate acid, conjugate base, Lewis concept of acid and base, Acid strength and pK_a .

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This note covers the following topics: Bronsted / Lowry Acid, Lewis Acid, Acids in Aqueous Solution, Relationship between K_b and K_a , Multiprotic acids, Equilibrium Calculations, Mass Balance, Equilibrium Relationship, Proton Condition, Weak Acid-Base Calculations, Strong Acid-Base Calculations, Strong Acid-Strong Base Mixtures, Strong acid-weak base mixtures, pH Buffers and Buffer Intensity, Carbonic Acid/Carbonate system.

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