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Atomic Mass & Mass Number Atomic mass and mass number are not the same. Atomic mass is the atom's actual mass 1 atomic mass unit (amu) is equal to 1/12 of the mass of the carbon-12 atom a proton has a mass of 1.007 amu a neutron has a mass of about 1.008 amu an electron has a very small mass of 0.00055 amu

### **Chapter 4 Atoms and Elements - Annville-Cleona School**

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36 Guided Reading and Study Workbook The Atomic Nucleus (pages 106–108) 15. Is the following sentence true or false? An alpha particle has a double positive charge because it is a helium atom that has lost two electrons.

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Bohr's contribution to the development of atomic structure a. was referred to as the 'plum pudding model'. b. was the discovery that electrons surround a dense nucleus. c. was proposed that electrons travel in circular orbits around the nucleus. d. is the quantum mechanical model. \_\_\_\_\_ 12.

### **Name Date Class ELECTRONS IN ATOMS 5**

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d. the atomic mass of the element \_\_\_\_\_ 16. What is the atomic number and the mass number of an atom with 11 protons and 12 neutrons? a. atomic number 11 and mass number 12 b. atomic number 12 and mass number 11 c. atomic number 11 and mass number 23 d. atomic number 23 and mass number 12 \_\_\_\_\_ 17. An atom of hydrogen-2 contains

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Isotopes of an element have the same atomic number but different mass numbers because they have different numbers of neutrons. Isotopes are atoms of the same element that have different numbers of neutrons and different mass numbers. It is hard to notice any differences in the properties of different isotopes of an element. 4.3 Modern ...

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Students use models to describe the atomic structure of a water molecule in "Modeling Atoms and Molecules" on page 13 of the TLR Introduction to Chemistry. They use models to describe the atomic composition of DNA in "Modeling the Genetic Code" on page 102 of the TLR Cells and Heredity.

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describe atomic structure in terms of number of protons, neutrons and electrons in the atom state the relative masses and relative charges of protons, neutrons and electrons represent isotopes using atomic numbers

and mass numbers, and use these together with ionic charges to calculate the number of protons, neutrons and electrons in atoms and ions

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Dalton's atomic theory stated that every element was made of atoms that could not be subdivided, atoms of the same element are alike, and a. atoms are made of protons, neutrons, and electrons.

### **CHAPTER 4 TEST: Atoms, Atomic Theory and Atomic Structure**

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2.2 Describe an atom and how atomic structure affects the mass number and atomic weight of the various chemical elements. 2.3 Explain the relationship between electrons and energy levels. 2.4 Compare the ways in which atoms combine to form molecules and compounds. 2.5 Describe the three states of matter and the

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When electrons are in the lowest energy state, they are said to be in the . ground state. When energy from a flame or other source is absorbed by the electrons, they are promoted to a

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4 ATOMIC STRUCTURE NOTES \_\_\_/\_\_\_ pts First Last Name \_\_\_ Per \_\_\_ Anticipatory Response Cornell Question & Ans 1. Matter can be divided? Y N Directions: Make 5 questions 2.

### **1 ATOMIC STRUCTURE NOTES / pts Last Name Per**

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Summarizing As you read, complete the table about atomic models. For more information on this Reading Strategy, see the Reading and Study Skills in the Skills and Reference Handbook at the end of

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ANALYSES AND CONCLUSIONS 1. How does this game simulate early efforts to determine the structure of the atom? In what ways is it different? 2. You eventually had the satisfaction of seeing the shape under the board.

### **ATOMIC STRUCTURE: 5 RUTHERFORD'S EXPERIMENT**

Atomic Number  $\neq$  Thus, the number of electrons (negatively charged particles) must equal the number of protons (positively charged particles).

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